High $\delta^{34}S$ in pyrite from magnetite-rich beds in the Urbana Limestone (eastern Sierra Morena).

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ABSTRACT

Anomalous high $\delta^{34}S$ (+39.7 o/oo and +45.5 o/oo) in pyrites from iron-beds in the Urbana Limestone are interpreted as a result of biogenic reduction of pore water sulphate in a system closed to $H_2S$ and $SO_4^{2-}$ and upwards migration of the remaining aqueous sulphate very depleted in $^{34}S$. The isotopic compositions of sulphur of pyrites from the Urbana Limestone, as well as those from Bancos Mixtos and Chavera Shales ($\delta^{34}S = +8.3$ o/oo and -13.6 o/oo, respectively) are consistent with the sedimentary environment assumed for these formations.

RESUMEN

Valores $\delta^{34}S$ anormalmente altos (+39.7 o/oo and 45.5 o/oo) observados en niveles ricos en magnetita y/o pirita en la Caliza Urbana son interpretados como el resultado de la reducción biogénica del sulfato en fase acusa en el sedimento, en un sistema cerrado a $H_2S$ y $SO_4^{2-}$ y una migración subsecuente del sulfato acusa remanente muy empobrecido en $^{34}S$. Los valores $\delta^{34}S$ observados en piritas de la Caliza Urbana, así como los de aquellas de los Bancos Mixtos y de las Pizarras Chavera son coherentes con el ambiente sedimentario interpretado para estas formaciones.

Key Words: sulphur isotopes, pyrite, magnetite, iron beds, Urbana Limestone.

Geogaceta, 18 (1995), 173-175
ISSN: 0213683X

Introduction

Pyrite found in pyrite- and magnetite-rich beds in the Urbana Limestone outcropping in eastern Sierra Morena displays anomalous high $\delta^{34}S$ values above the assumed $\delta^{34}S$ of the coeval Ordovician sea water (+27 o/oo after Claypool et al., 1980). Moreover, the occurrence of the heaviest $\delta^{34}S$ values at top of the set is inconsistent with the isotopic variations observed in modern sediments, where the light compositions are found close to the water-sediment interface (e.g. Vinogradov et al., 1962; Goldhaber and Klapan, 1980).

In the present contribution, we will examine the physico-chemical processes that could produce the observed isotopic anomalies.

Magnetite-pyrite beds in the Urbana Limestone

The studied section of the Urbana Limestone (early Ashgillian) is exposed at La Despreciada quarry, near Aldeaquemada (Jaén). It is made up of amalgamated sequences of decimetric and rare metric beds in accretional sets with reactivation surfaces and bimodal cross-laminae. The whole exposure (about 12 m. thick) is interpreted as nearshore tide-influenced clastic carbonate deposits (Lillo, 1992). A thin discontinuous bed of centimetric thickness formed by magnetite (dominant), pyrite, quartz (authigenic and detrital), siderite and minor berthieroid iron silicate is found in the lower portion of the exposure (Fig. 1). That iron bed is located a few cm. below a sedimentary reactivation surface. Magnetite and pyrite are mostly euhedral, pyrite in crystals up to one cm. size. Siderite is present as matrix. Berthieroid silicate appears in euhedral grains which can be deformed, surrounding the coarser grains of pyrite and magnetite. The presence of magnetite is limited to this bed but pyrite occurs (associated with accessory white mica) some cm. below, being relatively abundant in discontinuous lenses up to 3 cm. thick at 50 cm. below that ‘magnetite bed’.

Geochemical data

Two pyrite samples from the ‘pyrite bed’ and from the ‘magnetite bed’ were analysed for sulphur isotopes, yielding $\delta^{34}S = +39.7$ o/oo and $\delta^{34}S = +45.4$ o/oo

<table>
<thead>
<tr>
<th>Formation</th>
<th>$\delta^{34}S$ o/oo</th>
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<tbody>
<tr>
<td>Urbana Limestone</td>
<td></td>
</tr>
<tr>
<td>Magnetite bed</td>
<td>+ 45.4</td>
</tr>
<tr>
<td>Pirite bed</td>
<td>+ 39.7</td>
</tr>
<tr>
<td>Lumachelle</td>
<td>+ 16.5</td>
</tr>
<tr>
<td>Bancos Mixtos (Mixed Beds)</td>
<td></td>
</tr>
<tr>
<td>Chavera Shales</td>
<td>-13.6</td>
</tr>
</tbody>
</table>

Table 1. Sulphur Isotope data of sedimentary pyrites

Tabla 1.- Datos de isótopos de azufre de piritas sedimentarias

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or alternatively, it may precipitate at 'syngenetic-early diageneric' conditions as shown by Garrels and Christ (1965) (Fig. 2). In most of the ironstones, magnetite is found in mineral equilibrium with siderite and/or iron silicate (e.g. berthierite) and a syngenetic-early diageneric origin is commonly accepted if there are not unambiguous evidences to support the metamorphic origin (Brown, 1943). Similarly, pyrite, siderite and iron silicate can be stable at conditions encountered within the sediment (Curtis and Spears, 1968).

Pyrite formation in sediments is usually early diageneric, as a result of the reaction of $H_2S$ generated by bacterial reduction of the sulphate in pore water with reactive iron from minerals present in the sediment (Berner, 1970) (Fig. 2). In 'normal' marine sediments (those deposited in oxygenated bottom waters), the bacterial reduction of sulphate occurs in anoxic conditions, usually some cm below the water-sediment interface.

The bacteriogenic sulphide subsequently reacts with the iron present as $Fe^{2+}$ species (formed in post-oxic reactions) to produce metastable iron monosulphides that transform to pyrite (Fig. 2). In a closed system to $SO_2^-$ (which is the environment expected in the anoxic zone of 'normal' nearshore marine sediments), the remaining sulphate in pore water will be progressively depleted in $^{34}S$ (Goldhaber and Kaplan, 1980; Raiswell, 1982 among others). When the supply of iron is limited, the system is open to $H_2S$ as it remains in the system as aqueous specie. Alternatively, when an excess of reactive iron exists relative to the availability of aqueous sulphide (e.g. when the bacterial activity is limited by low amounts of readily metabolisable organic matter), the system is closed to $H_2S$: the aqueous sulphide is continuously removed to form pyrite. In that case, the sequence of pyrite deposition reflects a trend towards heavier $^{34}S$ values. Theoretically, if a Rayleigh distillation process is assumed, the $^{34}S$ value of sulphur incorporated in the latest formed pyrites may be higher than the original values of sulphate in sea water (Fig. 3).

Then, a previous $^{34}S$ depletion of $SO_2^-$ in pore water is required to explain the observed values. Note in fig. 3 that the $^{34}S$ values from pyrites formed in open systems to $H_2S$ are expected to be lower than those from closed systems and the $^{34}S$ value of sea water. Therefore, sedimentary pyrites...
with anomalous high δ³⁴S values above the δ³⁴S values of the coeval sea water (as the pyrites from the magnetite-pyrite beds in the Urbana Fm.) formed in an anoxic environment, a system closed to SO₄²⁻ and H₂S, during the early diagenesis of a 'normal' marine sediment.

If the sedimentary reactivation surface reflects a stage of winnowing or low rate of sediment deposition, then the occurrence of the heaviest isotopic values at top of the set is inconsistent with the variations of δ³⁴S values reported from modern sediments, where the light compositions are found close to the water-sediment interface (e.g. Vinogradov et al., 1962; Goldhaber and Kaplan, 1980). The stratigraphical 'inversion' of the δ³⁴S values in this section of the Urbana Limestone Fm. can be understood if the nature of these shallow-water sediments is considered. They deposited in a tidal nearshore environment, where the rate of sedimentation is high enough to produce a rapid compaction and subsequent vertical-upwards movement of the pore water in a well-sorted and relatively homogeneous sediment. This process leads to the relocation of the 'heavy' pore water (generated in the anoxic zone during the early stage of low rate of sedimentation) at higher stratigraphic levels; the formeroxic zone becomes rapidly an anoxic, sulphate depleted, iron-rich zone where most of the original organic matter has been already used in oxic reactions. In that environment the formation of sulphide is limited due to the low concentrations of dissolved sulphate and low amount of reactive organic matter, and siderite precipitates (cf. Spears, 1989). The later formation of siderite is supported by textual evidence.

The apparent large and inverse isotopic fractionation does not occur at large scale in the whole Urbana formation as it is indicated by the δ³⁴S value (+16.5 o/oo) of pyrite from a "lumachelle" in El Centenillo. That δ³⁴S value is reflecting a system closed to SO₄²⁻ but probably open to H₂S with a limited supply of iron. Similar conditions may be assessed for the formation of pyrite in the analysed sample from Bancos Mixtos, which yielded δ³⁴S = +8.2 o/oo, as they are also nearshore sediments (storm deposits close to wave base, after Lillo, 1992). Bacterial reduction of sea water sulphate in a system closed to SO₄²⁻ produces isotopic displacements up to 25 o/oo from the original values of seawater sulphate (Ohimoto and Rye, 1979; Goldhaber and Kaplan, 1980). Thus, the observed δ³⁴S values in pyrites from the Urbana Limestone and the Bancos Mixtos are very high for sulphide produced from marine sulphate if compared with the δ³⁴S values observed in pyrites formed in anoxic environments, where a slow rate of sulphate reduction produces large fractionation by kinetic effects, giving as result displacements in the δ³⁴S values between -40 o/oo and -60 o/oo (Ohimoto and Rye, 1979; Goldhaber and Kaplan, 1980). On this basis, the δ³⁴S py = -13.6 o/oo yielded by the sample from the Chavera Shales is consistent with pyrite formation in a relatively anoxic environment as expected for shelf muds.

Acknowledgments

I thank S. Bottrell for his fruitful comments. This work was carried out while at the Department of Earth Sciences (Univ. of Leeds). It was financed by a Spanish Ministry of Education grant.

References